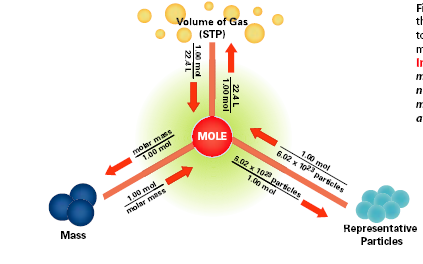
Unit Six Test Ch 13.1 p 385-389 10.2 p 300-303 and Name:

Ch 14.1 to 14.4(Dalton’s Law) p 413-433

Kinetic theory, kinetic energy, assumptions about gases, gas pressure, vacuum, atmospheric pressure, barometer, STP(p 300-303), pascals, atmospheres, mm of Hg, average kinetic energy, Kelvin temperature, Avogadro’s hypothesis, molar volume, molar volume at STP, n=v/V and n=m/M and v/V=m/M, (omit gas density calculations...replaced by conversions from mass🡨🡪volume. , Gas Stoichiometry, mol road map, compressibility, gas pressure-factors that affect it, amount of gas, volume, temperature, Gas Laws, Boyle’s Law, Charles’s Law,

Gay-Lussac’s Law, Combined Gas Law, Ideal Gas Law, Ideal Gas Constant, How are ideal gases and real gases different from one another?, Dalton’s Law of Partial Pressures, Ideal Gas Law @ two sets of

conditions. p 407 Do 26 to 35, p 315 Do 58e,59c,60,62 , p 439 39 to 59



Can you explain this **diagram?**

****

🡨omit

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